



Bal Bharati Public School, Pitampura

CLASS X CHEMISTRY ASSIGNMENT

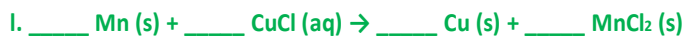
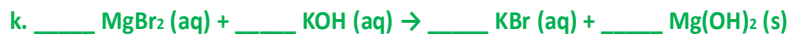
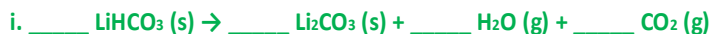
CHEMICAL REACTIONS AND EQUATIONS

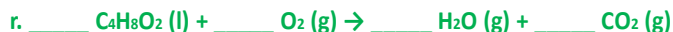
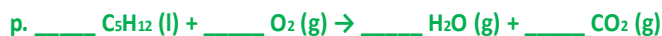
TERM I

Q1 Define:

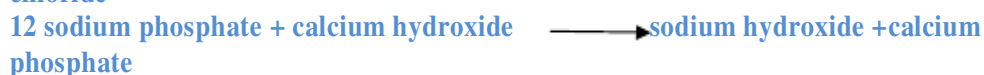
- Reactants
- Products
- Balanced chemical equation.

Q2 Balance the following chemical equation and identify the reactants and products.

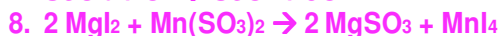
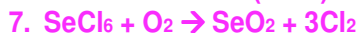
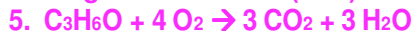
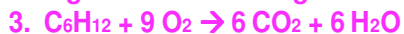
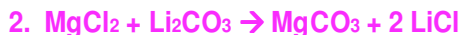




Q2 Write the skeleton equation for each of the following reactions. Then balance each of the following chemical equations.



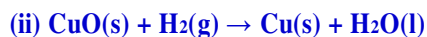
Q3 Indicate which type of chemical reaction (Combination, decomposition, single-displacement, double-displacement or combustion) is being represented in the following reactions:



Q4 Differentiate between:

- **Combination and Decomposition reaction.**
- **Displacement and Double displacement reaction.**
- **Oxidation and Reduction reaction.**
- **Balanced and unbalanced chemical equation.**
- **Corrosion of iron and Copper.**

Q5 Identify the substances that are oxidized and the substances that are reduced in the following reactions.



Study the above reaction and name the following:

- Substance getting reduced \rightarrow
- Substance getting oxidized
- Oxidizing agent
- Reducing agent.

Q7. Give reasons for the following:

- **Silver Nitrate solution cannot be stored in Copper containers.**
- **Gold and Silver do not corrode in air**
- **Blue colour of copper sulphate solution starts fading when a zinc rod is dipped in it.**
- **Respiration is an endothermic reaction.**
- **Photo chemical decomposition reaction finds application in photography.**

Q8. Why does stale food give a bad smell and bad taste? How can this be prevented?

Q9. For each of the following experiments, decide whether a reaction will occur. If you think there will be no reaction, write down why you think this.

If you think a reaction will happen:

write down why you think it will happen, what you expect to see, and the word equation to go with it.

- a) iron heated with copper(II) oxide
- b) aluminium heated with iron oxide
- c) copper heated with iron oxide
- d) magnesium heated with zinc oxide

Q10. Give examples for the following:

- **Precipitation reaction.**
- **Thermal decomposition.**
- **Natural oxidation.**
- **Exothermic reaction.**